PAPER SOLUTION

From Meerut

JEE MAIN 2026

JAN	SHIFT
22	1st

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JEE MAIN 2026 LIVE PAPER DISCUSSION

#Q. Match the following and choose the correct option.

	List-I		List-II
(a)	$[Ag(NH_3)_2]^+$	(i)	Fehling's solution
(b)	Zn-Hg/HCl	(ii)	Clemmenson's reduction
(c)	$NH_2 - NH_2/KOH$	(iii)	Tollen's reagent
(d)	Cu^{2+}/OH^-	(iv)	Wolff-Kishner reduction

A a–(i), b–(ii), c–(iii), d–(iv)

B a–(iv), b–(iii), c–(ii), d–(i)

C a–(iii), b–(ii), c–(iv), d–(i)

D a–(i), b–(ii), c–(iv), d–(iii)

Ans. (C)

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#Q. Statement-I : Sucrose is dextrorotary and upon hydrolysis it becomes laevorotatory.
Statement-II : Sucrose on hydrolysis gives glucose and fructose such that the laevorotation of glucose is more than dextrorotation of fructose.

(A) Both Statement-I and Statement-II are correct

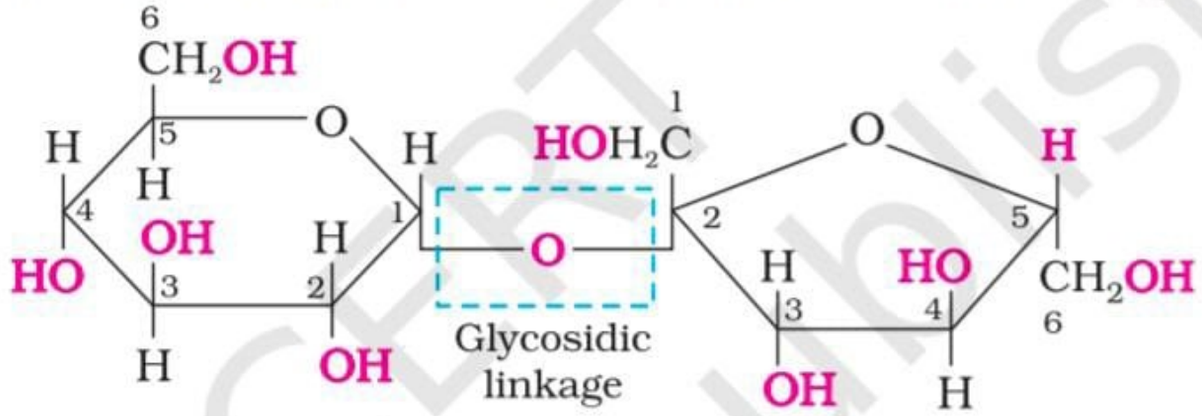
(B) Both Statement-I and Statement-II are incorrect

(C) Statement-I is correct, Statement-II is incorrect

(D) Statement-II is correct, Statement-I is incorrect

Ans. (C)

These two monosaccharides are held together by a glycosidic linkage between C1 of α-D-glucose and C2 of β-D-fructose. Since the reducing groups of glucose and fructose are involved in glycosidic bond formation, sucrose is a non reducing sugar.



α – D – Glucose β – D – Fructose

Sucrose

Sucrose is dextrorotatory but after hydrolysis gives dextrorotatory glucose and laevorotatory fructose. Since the laevorotation of fructose (–92.4°) is more than dextrorotation of glucose (+ 52.5°), the mixture is laevorotatory. Thus, hydrolysis of sucrose brings about a change in the sign of rotation, from dextro (+) to laevo (–) and the product is named as **invert sugar**.

#Q. Which of the following is the correct order of the reactivity of given nucleophiles when treated with CH_3Br in methanol?
F^-, I^-, $C_2H_5O^-$, $C_6H_5O^-$

(A) $I^- > C_2H_5O^- > C_6H_5O^- > F^-$

(B) $I^- > F^- > C_2H_5O^- > C_6H_5O^-$

(C) $I^- > C_2H_5O^- > F^- > C_6H_5O^-$

(D) $C_6H_5O^- > F^- > I^- > C_2H_5O^-$

Ans. (A)

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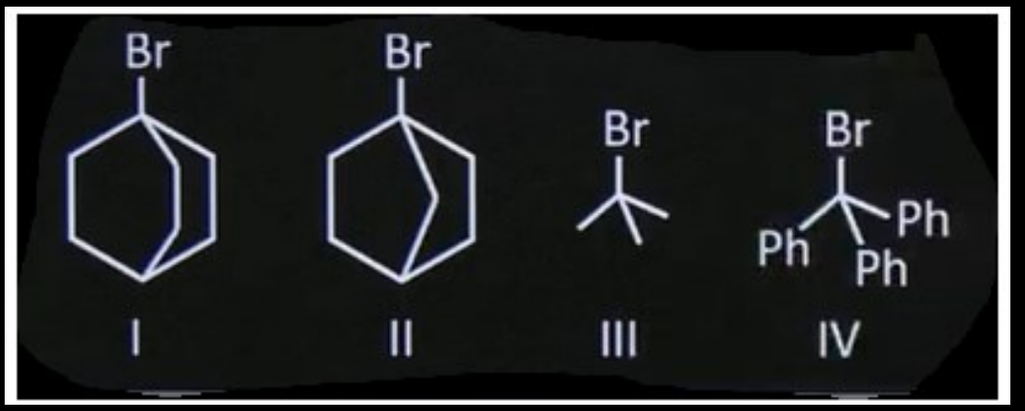
#Q. Reactivity of following on the basis S_N1 mechanism.



(A) IV > III > I > II

(B) II > IV > III > I

(C) III > IV > I > II

(D) IV > III > II > I

Ans. (A)

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#Q. Given below are two statements.
Statement I : HX bond length is higher in HCl than HF.
Statement II : PH_3 have the lowest boiling point in hydride of group 15 element.

A. Both Statement-I and Statement-II are correct

B. Both Statement-I and Statement-II are incorrect

C. Statement-I is correct, Statement-II is incorrect

D. Statement-II is correct, Statement-I is incorrect

Ans. (A)

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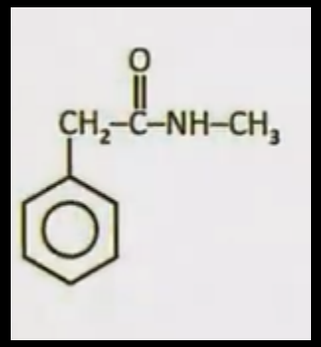
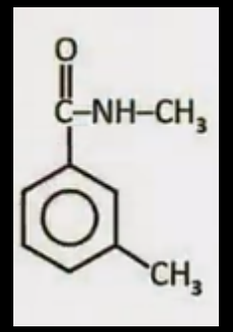
#Q. A compound 'A' with molecular formula C9H11NO reacts with Br2/NaOH at give (X). (X) on reaction with $NaNO_2$ in dil. HCl gives compounds (Y). is treated with CuCN, followed by hydrolysis gives (Z). The compound (A) on hydrolysis also gives compound (Z). Identify compound (A)

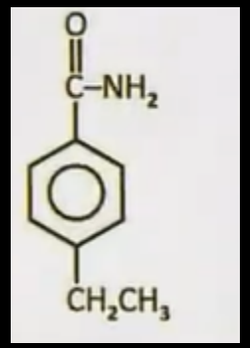
A



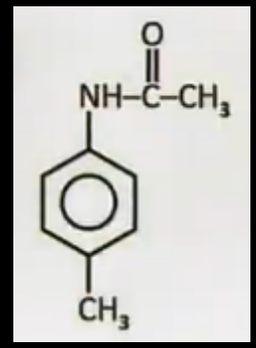
B



C



D



Ans. (C)

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#Q. Which of the following statement is correct regarding the nature and directive influence of $-NO_2$ group in nitration of benzene.

(A) It is an activating group and ortho/para director

(B) It is a deactivating group and ortho/para director

(C) It is a deactivating group and meta director

(D) It is an activating group and meta director

Ans. (C)

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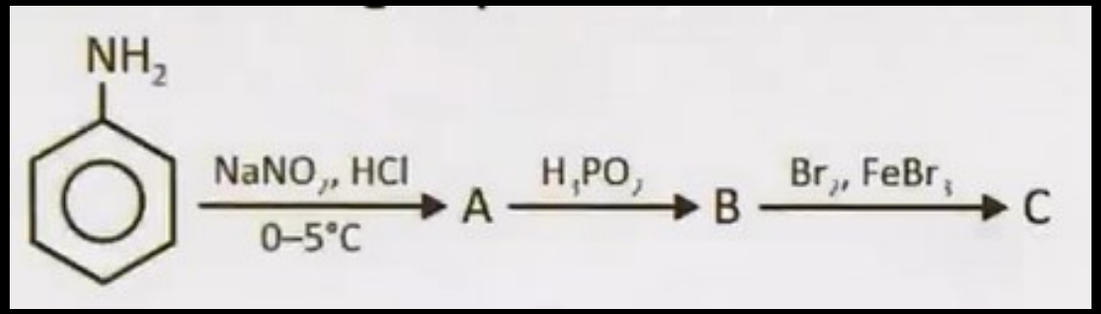
#Q. Consider the following sequence of reaction and identify A, B and C respectively.



(A) C_6H_5OH, C_6H_6, $C_6H_4Br_2$

(B) $C_6H_5N_2^+Cl^-$, C_6H_6, C_6H_5Br

(C) $C_6H_5NO_2$, C_6H_5OH, C_6H_5Br

(D) C_0H_5Cl, C_6H_5OH, C_6H_6

Ans. (B)

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#Q. Given below are two statements

Statement-I : K_H is constant with change in concentration of gas till solution is dilute at given temperature.

Statement-II : According to Henry's Law, partial pressure of gas in vapour phase is inversely proportional to mole fraction of gas in solution.

(A) Both Statement-I and Statement-II are correct

(B) Both Statement-I and Statement-II are incorrect

(C) Statement-I is correct, Statement-II is incorrect

(D) Statement-II is correct, Statement-I is incorrect

Ans. (C)

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#Q. Consider a first order reaction:

A → products

3 different solutions are taken rate of reaction

Solution 1 : 100mL 10M 'A' $-r_1$

Solution 2 : 200mL 10M 'A' $-r_2$

Solution 3 : 100mL 10M 'A' + 100mL water- r_3

The correct order of the rates of reactions is,

(A) $r_1 = r_2 = r_3$

(B) $r_1 = r_2 < r_3$

(C) $r_1 = r_2 > r_3$

(D) $r_1 < r_2 = r_3$

Ans. (C)

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#Q. Bohr's radius of H-atom is 2.12×10^{-10} m. Calculate the energy at this level.

(A) -5.44×10^{-19} J

(B) -2.176×10^{-18} J

(C) -54.4×10^{-19} J

(D) -2.3×10^{-19} J

Ans. (A)

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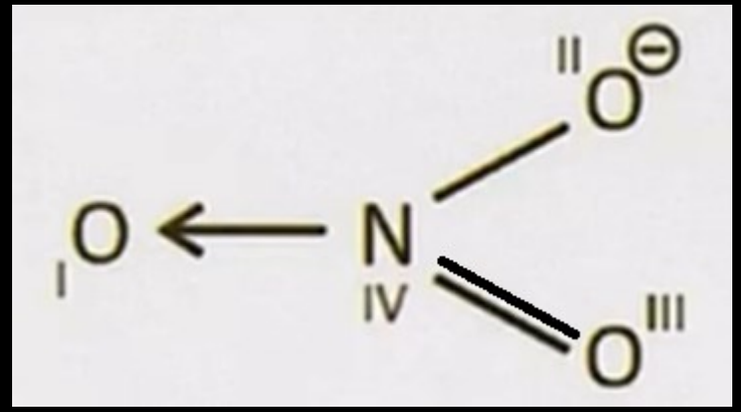
#Q. Find the formal charge of $\overset{IV}{N}$ and $\overset{I}{O'}$ $\overset{II}{O'}$ $\overset{III}{O}$.



(A) 0, +1, –1, +2

(B) +1, –1, –1, 0

(C) –1, 0, +2, +1

(D) +1, –1, 0, –1

Ans. (B)

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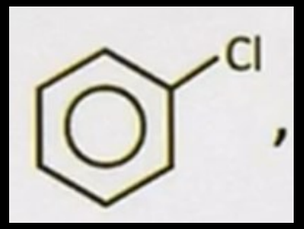
#Q. For ‘P’ 

, the incorrect statement is.

(A) In 'P' C-Cl bond has partial double bond character

(B) 'Cl' is an ortho-para directing group towards electrophilic aromatic substitution

(C) 'P' can undergo nucleophilic substitution reaction at normal conditions

(D) 'P' is less reactive than benzyl chloride towards nucleophilic substitution reaction.

Ans. (D)

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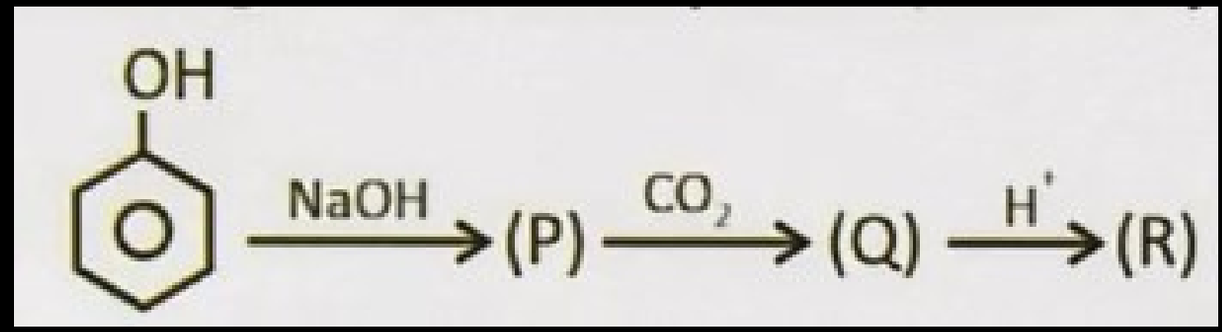
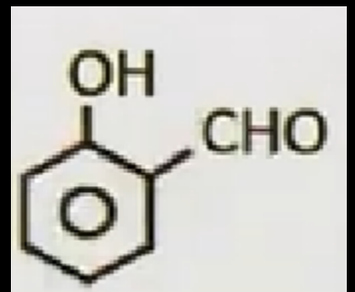
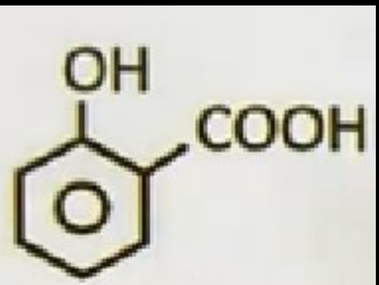
#Q. In the following reaction sequence, identify compound (R).



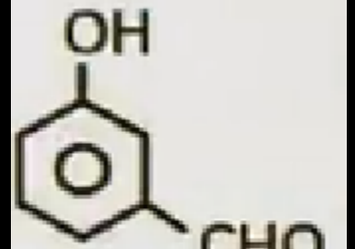
A



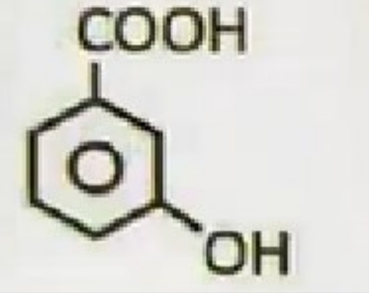
B



C



D



Ans. (A)

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#Q. Which of the following statement is correct regarding element having atomic number 79.

A It's first ionisation enthalpy is maximum in its group

B It's first ionisation enthalpy is minimum in its group

C It belongs to group 10 of periodic table

D It belongs to 5^{th} period of periodic table

Ans. (A)

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#Q. An element from 1St transition series and another element of 3rd transition series (same group) do not liberate H_2 gas from dilute acids like HCl. Both form halides. The hybridisation state of meatal ion halide respectively are:

(A) Both sp^3

(B) Both dsp^2

(C) sp^3 and dsp^2

(D) dsp^2 and sp^3

Ans. (C)

#Q. Sodium extract of organic compound of 0.1 g is treated with chlorin water and CCl_4 which dissolves in organic solvent produce a violet colour upon treatment with $AgNO_3$ a yellow ppt of 0.12 g is produce.
Calculate the percentage of Halide in organic compound.

Ans. (65)

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#Q. For the reaction given below at 25°C

$A_2 \leftrightharpoons 2A$

Find In K_p

Given $(\Delta G^\circ_f)_A$ = –50.384 KJ/ mol

Given $(\Delta G^\circ_f)_{S_2}$ = –100 KJ/ mol

A 0.43

B 0.23

C 0.31

D 0.53

Ans. (C)